



Roll No. of Candidate

Name of Candidate

Test Code: CT-04

CRASH TEST (MDCAT) SESSION-2021

Total MCQs: 120

Time Allowed: 120 min

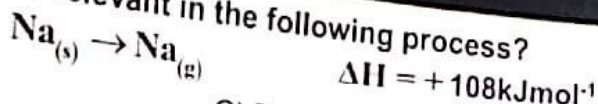
## BIOLOGY

- Phosphorylation of ADP during photosynthesis is called:**  
A) Oxidative photophosphorylation  
B) Photolysis  
C) Z scheme  
D) Oxidative phosphorylation
- Which reaction occurs only at night:**  
A) Dark Reaction  
B) Non-cyclic photophosphorylation  
C) Respiration  
D) None
- All are reactants of light reaction of photosynthesis except:**  
A) Blue Light  
B) Water  
C) Nicotinamide Adenine Dinucleotide  
D) Adenosine Diphosphate
- Decarboxylation occurs in all reactions except:**  
A) Alcoholic fermentation  
B) Pyruvic acid oxidation  
C) Krebs cycle  
D) Glycolysis
- How many moles of carbon dioxide are produced by the complete oxidation of 1 mole of pyruvate?**  
A) 1  
B) 3  
C) 4  
D) 6
- The reaction which occurs in thylakoid interior space:**  
A) Photolysis  
B) Transport of electrons  
C) ATP synthesis  
D) Dark reaction
- In which of the following steps, NADH is formed without decarboxylation?**  
A) Isocitrate –  $\alpha$ -ketoglutarate  
B)  $\alpha$ -ketoglutarate – succinate  
C) Pyruvate – Acetyl CoA  
D) Malate – oxaloacetate
- First compound formed during C3 cycle contains how many carbon atoms?**  
A) 3  
B) 6  
C) 4  
D) 5
- Feedback inhibitor of pyruvate decarboxylase is?**  
A) Succinate  
B) Malonate  
C) NADPH  
D) NADH
- All are products of Calvin cycle except:**  
A) Triose  
B) Inorganic phosphate  
C) Water  
D) Carbon dioxide
- Calvin cycle is known as C3 cycle because of:**  
A) Triose  
B) 3-phosphoglycerate  
C) Pyruvate  
D) 3 Carbon dioxide
- Dark reaction of photosynthesis occurs in:**  
A) Chloroplast  
B) Grana  
C) Mitochondria  
D) Stroma
- Source of protons for formation of NADPH in chloroplasts is:**  
A) Water  
B) Excited chlorophyll  
C) FADH  
D) Rubisco
- Which of the following is not a requirement of carbon fixation during Light independent reaction of photosynthesis?**  
A) NADPH  
B) Ribulose biphosphate  
C) Carbon dioxide  
D) Rubisco

15. How many ATP molecules are required for formation of hexose  
 A) 6  
 B) 9  
 C) 12  
 D) 18
16. Cellular respiration is:  
 A) Oxidation process  
 B) Reduction Process  
 C) Redox process  
 D) only Aerobic process
17. Which process occurs in mitochondria:  
 A) Alcoholic Fermentation  
 B) Glycolysis  
 C) Calvin cycle  
 D) Pyruvic acid oxidation
18. Phosphoenol pyruvate is formed through:  
 A) Decarboxylation  
 B) Dehydrogenation  
 C) Phosphorylation  
 D) Dehydration
19. Net ATP production during glycolysis and pyruvic acid oxidation respectively.  
 A) 4, 2  
 B) 2, 0  
 C) 2, 2  
 D) 10, 0
20. The net gain of energy from one molecule of glucose during aerobic respiration in prokaryotes is:  
 A) 2 ATP  
 B) 4 ATP  
 C) 38 ATP  
 D) 40 ATP
21. When a molecule of pyruvic acid is subjected to anaerobic respiration and forms lactic acid, there is:  
 A) 2 ATP are formed  
 B) 4 ATP are formed  
 C) No ATP is formed  
 D) 6 ATP are formed
22. Area of high concentration of proton in mitochondria is  
 A) Inter membrane space  
 B) Stroma  
 C) Matrix  
 D) Thylakoid interior space
23. How many molecules of ATP are required to convert a molecule of RuP into RuBP during Calvin cycle?  
 A) 1  
 B) 3  
 C) 2  
 D) 6
24. Each photon of light excite how many electrons?  
 A) 1  
 B) 2  
 C) 4  
 D) 6
25. During cyclic phosphorylation electron passes from all except:  
 A) Photosystem I  
 B) Photosystem II  
 C) Ferridoxin  
 D) Cytochrome complex
26. In yeast pyruvic acid is converted to?  
 A) Ethyl alcohol  
 B) Acetic acid  
 C) Lactic acid  
 D) Fumaric acid
27. The end product of the dark reaction is:  
 A) ATP  
 B) PGAL  
 C) RuBP  
 D) PEP
28. The final acceptor of electrons during oxidative phosphorylation is  
 A) Oxygen  
 B) ATP  
 C) Photosystem II  
 D) NADP<sup>+</sup>
29. Common product of alcoholic and lactic acid fermentation is  
 A) Carbon dioxide  
 B) lactate  
 C) Pyruvate  
 D) Alcohol
30. About 2% energy of chemical bonds of glucose is converted into ATP by:  
 A) Aerobic Respiration  
 B) Fermentation  
 C) Chemiosmosis  
 D) Calvin cycle



Which enthalpy change is relevant in the following process?



32. A) Enthalpy of fusion  
 B) Enthalpy of atomization  
 C) Enthalpy of vaporization  
 D) Enthalpy of formation

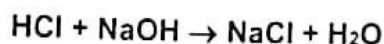
When two moles of  $\text{H}_2$  and one mole of  $\text{O}_2$  react to form  $\text{H}_2\text{O}$  484KJ heat is evolved what is  $\Delta H_f$  for one mole of  $\text{H}_2\text{O}$

33. A)  $-484 \text{ kJ mol}^{-1}$   
 B)  $-121 \text{ kJ mol}^{-1}$   
 C)  $-242 \text{ kJ mol}^{-1}$   
 D)  $+242 \text{ kJ mol}^{-1}$

For a given reaction  $\text{CH}_3\text{COOH} + \text{NaOH} \rightarrow \text{CH}_3\text{COONa} + \text{H}_2\text{O}$  the change in enthalpy under standard conditions is called?

34. A) Standard enthalpy change of solution  
 B) Standard enthalpy of neutralization  
 C) Standard enthalpy change of hydration  
 D) Standard enthalpy change of formation

Which of the equations shows the same "twice" the enthalpy change of neutralization as the following equation?



35. A)  $\text{H}_2\text{SO}_4 + \text{Mg}(\text{OH})_2 \rightarrow \text{MgSO}_4 + 2\text{H}_2\text{O}$   
 B)  $\text{MgCO}_3 + 2\text{HCl} \rightarrow \text{MgCl}_2 + \text{CO}_2 + \text{H}_2\text{O}$   
 C)  $\text{NH}_4\text{Cl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O} + \text{NH}_3$   
 D)  $\text{KOH} + \text{HCl} \rightarrow \text{KCl} + \text{H}_2\text{O}$

Lattice energy of an ionic crystal is the enthalpy of:

36. A) Combustion  
 B) Dissolution  
 C) Dissociation  
 D) Formation

In  $\text{MgCl}_2$ , the oxidation state of Cl is

37. A) Zero  
 B) +2  
 C) -2  
 D) -1

Which one of the following behave as a redox reaction?

38. A)  $\text{NaCl} + \text{AgNO}_3 \rightarrow \text{NaNO}_3 + \text{AgCl}$   
 B)  $2\text{Cl}^- \rightarrow \text{Cl}_2 + 2\text{e}^-$   
 C)  $2\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$   
 D)  $\text{Na}^+ + 1\text{e}^- \rightarrow \text{Na}$

In  $\text{SO}_4^{2-}$  the oxidation number of sulphur is

39. A) -8  
 B) +8  
 C) -6  
 D) +6

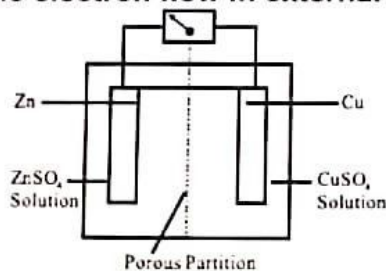
Study the following redox reaction:



Which statement is true about this reaction?

40. A) Manganese is oxidized from +7 to +2.  
 B) Chloride ions are reduced from -1 to zero  
 C) Chlorine is reduced from zero to -1  
 D) Manganese is reduced from +7 to +2

In the figure given below, the electron flow in external circuit is from



41. A) Zinc to copper electrode  
 B) Copper to zinc electrode  
 C) Right to left  
 D) porous partition to zinc electrode

The standard reduction potential of two electrodes is given as:

A = +1.36V, B = -0.44V the emf of the cell is:

42. A) +1.36V  
 B) -1.36V  
 C) +0.92V  
 D) +1.80V

Smaller is the value of standard reduction potential of substance:

- A) Greater is oxidizing power of substance  
 B) Greater is reducing power of substance  
 C) Lesser its tendency to combine with oxygen  
 D) Greater will be its tendency to displace the sodium from acid

43. Which of the following metals is most readily corroded in moist air?  
 A) Copper  
 B) Iron  
 C) Silver  
 D) Nickel
44. The metal which will act as anode when connected to copper:  
 A) Ag  
 B) Hg  
 C) Au  
 D) Al
45. Incorrect statement about electrolytic cells is:  
 A) They store a lot of energy  
 B) Anode is positively charged here  
 C) They carry out a non-spontaneous redox process  
 D) Reduction takes place at cathode
46. The enthalpy of atomization of chlorine can be \_\_\_\_\_ kJ/mole  
 A) -270  
 B) -240  
 C) +121  
 D) -218
47. Which of the followings always has a constant value?  
 A) Enthalpy of formation  
 B) Enthalpy of solution  
 C) Enthalpy of neutralization  
 D) Enthalpy of combustion
48. The amount of heat evolved or absorbed at constant pressure is  
 A) Work done  
 B) Enthalpy change  
 C) Internal energy  
 D) Free energy
49. The branch of chemistry, which deals with the thermal energy change is  
 A) Hydrodynamics  
 B) Thermodynamics  
 C) Thermo chemistry  
 D) Reaction Kinetics
50. When the energy associated with formation of bond is more than breakage of bond then reaction is \_\_\_\_\_  
 A) Endothermic  
 B) Exothermic  
 C) Not feasible  
 D) Heat is neither evolved nor absorbed
51. Enthalpy of the reaction does not give us information about  
 A) Nature of activated complex  
 B) Evolution or absorption of heat  
 C) Relative Stability of Product  
 D) Relative stability of Reactant
52. Heat given to a system not only increases the internal energy of the system but also does mechanical work. This is in accordance with  
 A) First law of thermo dynamics  
 B) 2<sup>nd</sup> law of thermo dynamics  
 C) 1<sup>st</sup> law of thermochemistry  
 D) Hess's law
53. Enthalpy of Neutralization is merely the  
 A) Enthalpy of Combustion  
 B) Enthalpy of decomposition  
 C) Enthalpy of formation of water from hydrated ions  
 D) Enthalpy of atomization
54. Which of the following is an example of endothermic process?  
 A) Freezing of water to ice  
 B) Respiration  
 C) Evaporation  
 D) Combustion of coal
55. The reaction  $2K_2MnO_4 + Cl_2 \rightarrow 2KMnO_4 + 2KCl$  is  
 A) Chlorination of  $K_2MnO_4$   
 B) Neither oxidation nor reduction  
 C) Neutralisation reaction  
 D) Oxidation reduction reaction
56. In the conversion of  $Br_2$  to  $BrO_3^-$ , the oxidation state of bromine changes from  
 A) 2 to 5  
 B) 1 to 5  
 C) 0 to -3  
 D) 0 to 5
57. The +1 oxidation state of P is found in  
 A)  $H_3PO_3$   
 B)  $H_3PO_4$   
 C)  $H_3PO_2$   
 D)  $H_4P_2O_7$
58. Electrode potential is effected by \_\_\_\_\_  
 A) Molarity  
 B) Temperature  
 C) Pressure  
 D) Temperature, pressure and concentration of ions
59. The oxidation number of carbon in  $C_2H_6$  is  
 A) -3  
 B) -6  
 C) +6  
 D) +2
60. Electrolysis of the aqueous solution of which of the followings does not produce Oxygen gas  
 A)  $CaCl_2$   
 B)  $Pb(OH)_2$   
 C)  $NaNO_3$   
 D)  $CuSO_4$



The value of universal gas constant R is

- A) 8.314 J/K  
 B) 8.314 J mol<sup>-1</sup> K<sup>-1</sup>  
 C) 8314 J mol<sup>-1</sup> K<sup>-1</sup>  
 D) 83.14 J mol<sup>-1</sup> K<sup>-1</sup>

2. If  $C_v = \frac{5}{2} R$  then  $C_p$  in

- A)  $\frac{2}{5} R$   
 B)  $\frac{2R}{7}$   
 C)  $\frac{7}{2} R$   
 D)  $\frac{5}{2} R$

63. Which of the following is not a state function of a system?

- A) Temperature  
 B) Heat  
 C) Internal energy  
 D) Pressure

64. The equation of state for 5 g of oxygen at a pressure P and temperature T, when occupying a volume V, will be

- A)  $PV = 5 RT$   
 B)  $PV = \left(\frac{5}{16}\right) RT$   
 C)  $PV = \left(\frac{5}{2}\right) RT$   
 D)  $PV = \left(\frac{5}{32}\right) RT$

65. A gas behaves as an ideal gas:

- A) At low temperature and high-pressure  
 B) High pressure high temperature  
 C) At low pressure and high temperature  
 D) Low pressure low temperature

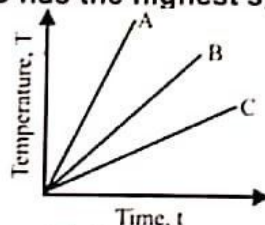
66. At the same temperature that mean kinetic energies of molecules of hydrogen and oxygen are in the ratio

- A) 1 : 1  
 B) 1 : 16  
 C) 8 : 1  
 D) 16 : 1

67. Four students found set of  $C_p$  and  $C_v$  (in cal/deg mole) as given below. Which of the following set is correct?

- A)  $C_v = 4, C_p = 2$   
 B)  $C_v = 2, C_p = 1$   
 C)  $C_v = 3, C_p = 3$   
 D)  $C_p = 5, C_v = 3$

68. Which of the substances A, B, or C has the highest specific heat?



- A) A  
 B) B  
 C) C  
 D) All have equal specific heat

69. The specific heat of a gas at constant pressure as compared to that at constant volume is

- A) Less  
 B) Equal  
 C) More  
 D) Constant

70. Two identical sample of a gas are allowed to expand (i) isothermally (ii) adiabatically. Work done is

- A) More in the isothermal process  
 B) More in the adiabatic process  
 C) Neither of them  
 D) Equal in both processes

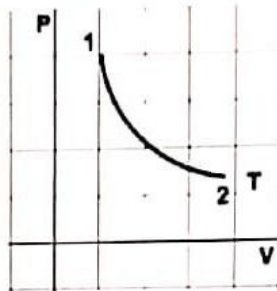
71. For hydrogen gas  $C_p - C_v = a$ , and for oxygen gas  $C_p - C_v = b$ , so that relation between a and b given by

- A)  $a = 16 b$   
 B)  $16 a = b$   
 C)  $a = b$   
 D)  $a = 4 b$

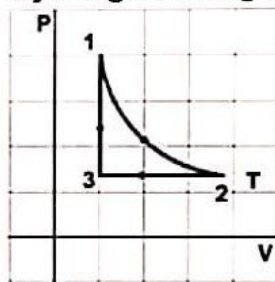
72. In an adiabatic process,  $PV^\gamma = \text{constant}$ ; the value ' $\gamma$ ' is

- A)  $\frac{C_p}{C_v}$   
 B)  $1 - \frac{R}{C_v}$   
 C)  $\frac{R}{C_v} - 1$   
 D)  $1 + \frac{R}{C_v}$

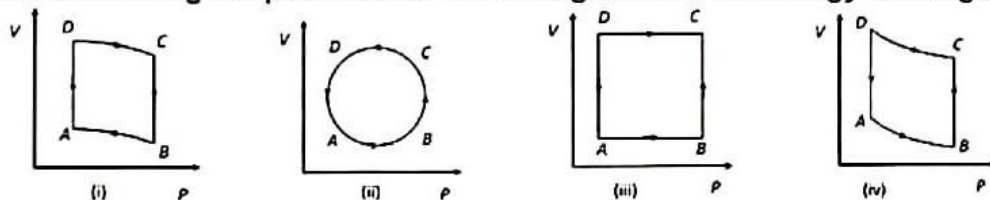
73. The state of an ideal gas is changed isothermally from position 1 to position 2 above. What is the change in the internal energy of the gas during this process?



- A)  $\Delta U = W$   
 B)  $\Delta U > 0$   
 C)  $\Delta U = Q$   
 D)  $\Delta U = 0$
74. A sample of an ideal gas taken through a closed cycle is presented by the P-V diagram. The process 1-2 is perfectly isothermal. Which of the following is true about the change in internal energy and work done by the gas during the process 1-2?



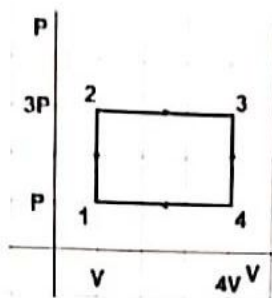
- A)  $\Delta U = 0$       W by the gas  $> 0$   
 B)  $\Delta U > 0$       W by the gas  $= 0$   
 C)  $\Delta U < 0$       W by the gas  $< 0$   
 D)  $\Delta U = 0$       W by the gas  $= 0$
75. Which is called internal energy of an ideal gas?  
 A) Potential energy  
 B) Vibrational kinetic energy  
 C) Translational kinetic energy  
 D) Rotational kinetic energy
76. In the diagrams (i) to (iv) of variation of volume with changing pressure is shown. A gas is taken along the path ABCD. The change in internal energy of the gas will be



- A) Positive in all cases (i) to (iv)  
 B) Positive in cases (i), (ii) and (iii) but zero in (iv) case  
 C) Negative in cases (i), (ii) and (iii) but zero in (iv) case  
 D) Zero in all four cases
77. Cloud formation in atmosphere is an example of  
 A) isothermal process  
 B) adiabatic process  
 C) isochoric process  
 D) isobaric process
78. The amount of heat energy required to raise the temperature of a body of mass 1 kg through 1 k is called:  
 A) Specific heat  
 B) Heat capacity  
 C) Molar specific heat  
 D) Heat of vaporization
79. At a constant temperature when the volume of a given mass of a gas is doubled, its density becomes:  
 A) Double  
 B) Four times  
 C) One fourth  
 D) Half



An ideal gas is taken through a closed path 1 → 2 → 3 → 4 → 1. What is the net work done by the gas?



- A) 6PV  
B) 4PV  
C) 9PV  
D) -6PV
81. The temperature of 2 mole of a gas is changed. From 100°C to 120°C at constant volume. The change in internal energy was found to be 80 J. What is the molar heat capacity of this gas at constant volume?  
A) 0.4 J K<sup>-1</sup> mol<sup>-1</sup>  
B) 4 J K<sup>-1</sup> mol<sup>-1</sup>  
C) 2.0 J K<sup>-1</sup> mol<sup>-1</sup>  
D) 8 J K<sup>-1</sup> mol<sup>-1</sup>
82. If 1 mole of an ideal gas is heated at constant pressure then  
A)  $Q_p = C_v \Delta T$   
B)  $Q_v = C_v \Delta T$   
C)  $Q_p = C_p \Delta T$   
D)  $Q_v = C_p \Delta T$
83. For 1 mole of gas the relation  $P\Delta V =$   
A)  $R\Delta T$   
B)  $R\Delta P$   
C)  $R\Delta V$   
D)  $P\Delta T$
84. A gas expands 0.25 m<sup>3</sup> at constant pressure  $10^5 \text{ N/m}^2$ , the work done is  
A) 2.5 ergs  
B) 250 W  
C) 250 J  
D) 250 N
85. If 20J of work is done in compressing a gas adiabatically the change in internal energy is equal to  
A) 20J  
B) -20J  
C) 10J  
D) 200J
86. In a certain process, 400J of heat energy is supplied to a system and at the same time 150 J of work is done by the system. What is the increase in internal energy of the system?  
A) 250 J  
B) 50 J  
C) 100 J  
D) 150 J
87. The PV diagrams representing maximum and minimum amount of work done are respectively  
  
A) A and B  
B) B and C  
C) B and D  
D) C and D
88. Internal energy of a system is defined as  
A) The sum of kinetic energies of all molecules of the system  
B) The sum of kinetic and potential energies of all molecules of the system  
C) The sum of potential energies of the system  
D) The average kinetic energy of all molecules
89. 110 J of heat is added to a gaseous system, whose internal energy change is 40 J, then the amount of external work done is  
A) 150 J  
B) 110 J  
C) 70 J  
D) 40 J
90. When compressed gas is suddenly allowed to expand, which of the following equation determines the P - V relationship with  $\gamma$  being the gas constant?  
A)  $PV = \gamma$   
B)  $VP^\gamma = \text{constant}$   
C)  $PV^\gamma = \text{constant}$   
D)  $PV/\gamma = \text{constant}$



## SPOT THE ERROR:

In the following sentences some segments of each sentence are underlined. Your task is to identify the underlined segment of the sentence, which contains the mistake that needs to be corrected.

91. Chips would sing out, in that jerky, high-pitch voice that had still a good deal of  
 A) B) C) D)
92. I remember your father - he used to sit at that far desk by the wall - he was not very better, either.  
 A) B) C) D)
93. Mr. Chips's room was furnished simply and with schoolmaster taste.  
 A) B) C) D)
94. I feel lonely and isolated because I have a few friends in this class.  
 A) B) C) D)
95. When we arrived at the bus station, it turned out that the latest bus had already left.  
 A) B) C) D)
96. The bright boy is the more pitiable object among all our failing students.  
 A) B) C) D)
97. They built a white small cottage and ate good breakfasts there.  
 A) B) C) D)
98. We seem to have strength, but it is only the appearance of strength; otherwise we  
 A) B) C) D)
- are weak and without motion.

In each of the following questions, four alternative sentences are given.

99.  
 A) The summer temperature in Los Angeles is much high than San Francisco.  
 B) The summer temperature in Los Angeles is much higher than that in San Francisco.  
 C) The summer temperature in Los Angeles is very higher than San Francisco.  
 D) The summer temperature in Los Angeles is much higher than San Francisco.
100.  
 A) I, therefore, had to throw them away or wipe them off the map altogether.  
 B) I therefore, had to throw them away or wipe them off the map all together.  
 C) I therefore, had to throw them away or wipe it off the map altogether.  
 D) I, therefore, had to throw them away or wipe them off the map all together.
101.  
 A) It would be hard to think of a less appetizer drink than the greenish fluid.  
 B) It would be hardly to think of a less appetizing drink than the greenish fluid.  
 C) It would be hard to think of a least appetizing drink than the greenish fluid.  
 D) It would be hard to think of a less appetizing drink than the greenish fluid.
102.  
 A) The parents of the gold medalist were much pleased at the success of their son.  
 B) The parents of the gold medalist were too pleased at the success of their son.  
 C) The parents of the gold medalist were pleased a lot at the success of their son.  
 D) The parents of the gold medalist were very pleased at the success of their son.
103.  
 A) He is a pleasant looking Man of about thirty five year.  
 B) He is a pleasant looking Man about thirty-five.  
 C) He is a pleasant looking Man of about thirty-five-years.  
 D) He is a pleasant looking Man of about thirty-five.
104.  
 A) Housman believes that this life is too short to live.  
 B) Housman believes that this life is much short to live.  
 C) Housman believes that this life is rather short to live.  
 D) Housman believes that this life is very short to live.
105.  
 A) Your condition is as better as theirs, but you are different from them in certain aspects.  
 B) Your condition is as good as theirs, but you are different from them in certain aspects.  
 C) Your condition is as better as theirs, but you are different than them in certain aspects.  
 D) Your condition is as good as that of theirs, but you are different from them in certain aspects.



- It is like that there are those who jump at the opportunity of getting rid of books they don't want.  
 A) B) C) D)
7. He was claiming her interest and attention farther more than youths of her own age.  
 A) B) C) D)
108. You will all be deep grieved to hear that His Majesty King Edward the Seventh died this morning  
 A) B) C) D)

**SYNONYMS:**

109. **GINGERLY**  
 A) Heedlessly  
 B) Daintily  
 C) Falsely  
 D) Recklessly
110. **HIDEOUS**  
 A) Resplendent  
 B) Dazzling  
 C) Brilliant  
 D) Obnoxious

**LOGICAL REASONING**

111.  
 I. Police had resorted to lathi-charge to disperse the unruly mob from the civic headquarters.  
 II. The civic administration has recently hiked the property tax of the residential buildings by about 30 percent.  
 A) Statement I is the cause and statement II is its effect.  
 B) Statement II is the cause and statement I is its effect.  
 C) Both the statements I and II are independent causes.  
 D) Both the statements I and II are effects of independent causes.  
 E) Both the statements I and II are effects of some common cause.
112.  
 I. The government has allowed private airline companies in India to operate to overseas destinations.  
 II. The national air carrier has increased its flights to overseas destinations.  
 A) Statement I is the cause and statement II is its effect.  
 B) Statement II is the cause and statement I is its effect.  
 C) Both the statements I and II are independent causes.  
 D) Both the statements I and II are effects of independent causes.  
 E) Both the statements I and II are effects of some common cause.
113.  
 I. Large number of people living in the low lying areas has been evacuated during the last few days to safer places.  
 II. The Government has rushed in relief supplies to the people living in the affected areas.  
 A) Statement I is the cause and statement II is its effect.  
 B) Statement II is the cause and statement I is its effect.  
 C) Both the statements I and II are independent causes.  
 D) Both the statements I and II are effects of independent causes.  
 E) Both the statements I and II are effects of some common cause.
114.  
 I. The performance of Indian sports persons in the recently held Olympics could not reach the level of expectation the country had on them.  
 II. The performance of Indian sports person in the last Asian games was far better than any previous games.  
 A) Statement I is the cause and statement II is its effect.  
 B) Statement II is the cause and statement I is its effect.  
 C) Both the statements I and II are independent causes.  
 D) Both the statements I and II are effects of independent causes.  
 E) Both the statements I and II are effects of some common cause.

115. I. Large numbers of people have fallen sick after consuming sweets from a particular shop in the locality.  
 II. Major part of the locality is flooded and has become inaccessible to outsiders.  
 A) Statement I is the cause and statement II is its effect.  
 B) Statement II is the cause and statement I is its effect.  
 C) Both the statements I and II are independent causes.  
 D) Both the statements I and II are effects of independent causes.  
 E) Both the statements I and II are effects of some common cause.
116. I. The life today is too fast, demanding all full of variety in all aspects which at times leads to stressful situation.  
 II. Number of suicide cases among teenagers is on increase.  
 A) Statement I is the cause and statement II is its effect.  
 B) Statement II is the cause and statement I is its effect.  
 C) Both the statements I and II are independent causes.  
 D) Both the statements I and II are effects of independent causes.  
 E) Both the statements I and II are effects of some common cause.
117. I. Majority of people residing in the locality have decided to protest against the municipality authority's decision to allow construction of a shopping mall in the locality.  
 II. Many shopping malls have been opened all over the city in the recent time.  
 A) Statement I is the cause and statement II is its effect.  
 B) Statement II is the cause and statement I is its effect.  
 C) Both the statements I and II are independent causes.  
 D) Both the statements I and II are effects of independent causes.  
 E) Both the statements I and II are effects of some common cause.
118. I. The Government has reduced the prices of petroleum products by five percent a week after increasing the prices by ten percent.  
 II. The rate of inflation dropped marginally during the last week.  
 A) Statement I is the cause and statement II is its effect.  
 B) Statement II is the cause and statement I is its effect.  
 C) Both the statements I and II are independent causes.  
 D) Both the statements I and II are effects of independent causes.  
 E) Both the statements I and II are effects of some common cause.
119. I. Large number of Primary Schools in the rural areas is run by only one teacher.  
 II. There has been a huge dropout from the primary schools in rural areas.  
 A) Statement I is the cause and statement II is its effect.  
 B) Statement II is the cause and statement I is its effect.  
 C) Both the statements I and II are independent causes.  
 D) Both the statements I and II are effects of independent causes.  
 E) Both the statements I and II are effects of some common cause.
120. I. The car manufacturing companies have recently increased the prices of mid-sized cars.  
 II. The Government recently increased the duty on mid-sized cars.  
 A) Statement I is the cause and statement II is its effect.  
 B) Statement II is the cause and statement I is its effect.  
 C) Both the statements I and II are independent causes.  
 D) Both the statements I and II are effects of independent causes.  
 E) Both the statements I and II are effects of some common cause.



# CRASH TEST SESSION 2021

1	<input checked="" type="radio"/> A	<input type="radio"/> B	<input checked="" type="radio"/> C	<input type="radio"/> D	31	<input type="radio"/> A	<input checked="" type="radio"/> B	<input type="radio"/> C	<input type="radio"/> D	61	<input type="radio"/> A	<input checked="" type="radio"/> B	<input type="radio"/> C	<input type="radio"/> D	91	<input type="radio"/> A	<input checked="" type="radio"/> B	<input type="radio"/> C	<input type="radio"/> D
2	<input type="radio"/> A	<input type="radio"/> B	<input type="radio"/> C	<input checked="" type="radio"/> D	32	<input type="radio"/> A	<input type="radio"/> B	<input checked="" type="radio"/> C	<input type="radio"/> D	62	<input type="radio"/> A	<input type="radio"/> B	<input checked="" type="radio"/> C	<input type="radio"/> D	92	<input type="radio"/> A	<input type="radio"/> B	<input checked="" type="radio"/> C	<input type="radio"/> D
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4	<input type="radio"/> A	<input type="radio"/> B	<input type="radio"/> C	<input checked="" type="radio"/> D	34	<input checked="" type="radio"/> A	<input type="radio"/> B	<input type="radio"/> C	<input type="radio"/> D	64	<input type="radio"/> A	<input type="radio"/> B	<input type="radio"/> C	<input checked="" type="radio"/> D	94	<input type="radio"/> A	<input type="radio"/> B	<input type="radio"/> C	<input checked="" type="radio"/> D
5	<input checked="" type="radio"/> A	<input checked="" type="radio"/> B	<input type="radio"/> C	<input type="radio"/> D	35	<input type="radio"/> A	<input type="radio"/> B	<input type="radio"/> C	<input checked="" type="radio"/> D	65	<input type="radio"/> A	<input type="radio"/> B	<input checked="" type="radio"/> C	<input type="radio"/> D	95	<input type="radio"/> A	<input type="radio"/> B	<input checked="" type="radio"/> C	<input type="radio"/> D
6	<input checked="" type="radio"/> A	<input type="radio"/> B	<input type="radio"/> C	<input checked="" type="radio"/> D	36	<input type="radio"/> A	<input type="radio"/> B	<input type="radio"/> C	<input checked="" type="radio"/> D	66	<input checked="" type="radio"/> A	<input type="radio"/> B	<input type="radio"/> C	<input type="radio"/> D	96	<input checked="" type="radio"/> A	<input type="radio"/> B	<input type="radio"/> C	<input type="radio"/> D
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