PMC paid practice test bundle 2 paper 5

Physics

Q 1

The root mean square value of the alternating current is equal to

A

twice the peak value

B

half the peak value

C

equal to the peak value

D

(1/√2) times the peak value

Q 2

What will happen in a time of 7 hours, if a radioactive substance has an average life of 7 hours?

A

Half of the active nuclei decay

B

less half of the active nuclei decay

C

more than half of the active nuclei decay

D

total nuclei decay

Q 3

The principle behind the working of cathode ray oscilloscope is

A

oscillation

B

half wave rectification

C

full wave rectification

D

none of these

Q 4

The magnitude of the displacement is:...

A

size of object A

B

straight line distance between the initial position and the final position of the body

C

size of object B

D

any distance between the initial position and the final position of the body

Q 5

In stationary waves

A

Strain is maximum at antinodes

B

Strain is minimum at nodes

C

Strain is maximum at node

D

Amplitude is same at all points

Q 6

When temperature of the system does not change, it is called:...

A

isothermal

B

adiabatic

C

isovolumetric

D

isobaric

Q 7

Which of the following types of force can do no work on the particle upon which it acts

A

frictional force

B

gravitational force

C

centripetal force

D

elastic force

Q 8

The negative sign in F = -k.x, indicates that:

A

F is directed opposite to x

B

F is directed along x

C

F is always equal to x

D

none of these

Q 9

Find the current if power given is 5 Watts and voltage is 0.5 volts.

A

10 Amp

B

20 Amp

C

30 Amp

D

50 Amp

Q 10

What does the constant N represent in the equation of state for an ideal gas PV=NkT?

A

number of molecules of gas

B

number of moles of the gas

C

number of nucleons

D

number of protons

Q 11

The waves which propagate by the oscillation of material particles are called:

A

matter waves

B

mechanical waves

C

electromagnetic waves

D

microwaves

Q 12

The body move 10 m along the straight line under the force 5 N, what is the angle between force and displacement

A

0 degree

B

30 degree

C

45 degree

D

60 degree

Q 13

Which of the following can have negative temperature coefficient?

A

Compounds of silver

B

Liquid metals

C

Metallic alloys

D

Electrolytes

Q 14

A line normal to the wavefront, showing the direction of propagation of light is called:

A

tangent

B

radius

C

wavelength

D

ray of light

Q 15

Range of wavelength of visible light is :

A

700˚A - 1000˚A

B

1nm - 100nm

C

0.1nm - 1nm

D

4000˚A - 7000˚A

Q 16

The phenomenon of radioactivity is

A

Nuclear process does not depend on external factors

B

increased on applied pressure

C

exothermic change which increase and decrease with temperature

D

none of these

Q 17

The SI unit of angular momentum is Js. It can also be expressed as:

A

kg ms^-1

B

kgm^2s^-1

C

kgm^2s^-2

D

kgm^-2-1

Q 18

A sound wave travels from a region of hot air into a region of cold air. How does frequency and wavelength of sound change?

A

frequency decreases wavelength decreases

B

frequency increases wavelength decreases

C

frequency does not change wavelength decreases

D

frequency does not change wavelength does not change

Q 19

The temperature coefficient of resistance is expressed in :

A

°C

B

°C-1

C

m°C-1

D

None of these

Q 20

Half wave rectifier passes only

A

lower half cycle

B

upper half cycle

C

both cycles

D

none of them

Q 21

Electric potential difference can be defined as:...

A

$\Delta V = \Delta U - q$

B

$\Delta V = \Delta U / q$

C

$\Delta V = \Delta U + q$

D

$\Delta V = \Delta U * q$

Q 22

Calculate the frequency if the number of revolutions is 300 and the paired poles are 50.

A

15 kHz

B

150 kHz

C

1500 kHz

D

150 Hz

Q 23

A transformer has negative voltage regulation when its load power factor is

A

Lagging

B

Leading

C

Unity

D

any of above

Q 24

The proton and antiproton collision will result

A

scattering

B

repulsion

C

attraction

D

annihilation

Q 25

A full wave rectifier passes _____into positive cycles

A

lower half cycle

B

upper half cycle

C

both cycles

D

none of them

Q 26

The magnetic field lines generated in current carrying conductor are

A

circular

B

triangular

C

linear

D

none of these

Q 27

To start a fusion reaction, energy required is

A

small

B

large

C

infinite

D

zero

Q 28

One kcal =

A

4.18 J

B

4180 J

C

2.09 J

D

2090 J

Q 29

An object moves 20 m in 5 sec. What is the gradient of the displacement-time graph?

A

25

B

15

C

4

D

1/4

Q 30

If a horse pulls a cart, work done by horse is

A

negative

B

zero

C

positive

D

none of these

Q 31

The units of angular velocity are similar to

A

angular displacement

B

angular acceleration

C

angular frequency

D

none of these

Q 32

Strength of magnetic field is called

A

strength

B

flux

C

magnetic flux density

D

density

Q 33

Ripple factor is defined as

A

Irms/Vrms

B

Idc/Irms

C

Irms/I dc

D

Irms + I dc

Q 34

An object is undergoing simple harmonic motion. It's time period is T and total energy is E. The amplitude of vibration is reduced to half. What is the new time period and total energy of the system?

A

time period = T/2 total energy = E/4

B

time period = T total energy = E/4

C

time period = T/2 total energy = E/2

D

time period = T total energy = E/2

Q 35

X-ray is the reverse process of :

A

Pair production

B

Compton effect

C

Photoelectric effect

D

A & B are correct

Q 36

Bright Lines in pattern shows

A

absorption

B

emission

C

release

D

free particle

Q 37

The angular speed of the wheels of a bicycle is 8π radian/sec there period of rotation is :

A

.25 sec

B

.4 sec

C

π/4 sec

D

4 sec

Q 38

Electric field at a point varies as r^0 for

A

a plane infinite sheet of charge

B

A point charge

C

Electric dipole

D

Line charge of infinite length

Q 39

The first law of thermodynamics can be stated as:

A

Q = ΔU + W

B

Q + ΔU = W

C

Q = ΔU - W

D

Q = ΔU * W

Q 40

Current and magnetic field are always

A

perpendicular

B

parallel

C

circular

D

none of these

Q 41

Req = R1 + R2 + R3 +..... + Rn is the combination in

A

Series

B

Parallel

C

Both of them

D

None of them

Q 42

Electric potential difference between the two points can be defined as:...

A

difference of the kinetic energy per unit charge

B

difference of the kinetic energy

C

difference of the potential energy per unit charge

D

difference of the potential energy

Q 43

Which of the following substances cannot be emitted by radioactive substances during their decay?

A

proton

B

neutrinos

C

helium nuclei

D

electrons

Q 44

Specific latent heat of fusion of ice is 334 J/g. How much energy is needed to melt 100 g of ice at 0C.

A

33.4 J

B

33.4 kJ

C

3.34 J

D

3.34 kJ

Q 45

How much energy is required to raise the temperature of 5.00 kg of lead from 20.0°C to its melting point of 327°C? The specific heat of lead is 128 J/kg ? °C.

A

1.96 x 10^5 J

B

4.04 x 10^5 J

C

1.07 x 10^5 J

D

8.15 x 10^4 J

Q 46

The shortest distance between two points on a travelling wave that have a phase difference of (pi/3) is 5 cm. If the wave has frequency 500 Hz, what is the speed of the wave?

A

300 m/s

B

150 m/s

C

300 cm/s

D

150 cm/s

Q 47

Energy stored in an inductor is

A

electric energy

B

magnetic energy

C

electromagnetic energy

D

both a) and b)

Q 48

100 W heater is used to melt 50 g of ice at 0C. How long should the heater be switched on? Specific latent heat of fusion of ice is 334 J/g.

A

28 min

B

16.7 min

C

2.8 min

D

167 min

Q 49

The isotope of 235U92 has _____ number of neutrons

A

141

B

142

C

143

D

144

Q 50

Calculate the charge passing through the circuit if it's current it 10 Amp and the recorded time is 15 seconds

A

1500 Coulomb

B

150 Coulomb

C

13400 Coulomb

D

140 Coulomb

Q 51

When gamma photon is entered in nucleus it_____

A

de-excite the atom

B

excite the atom

C

scatter by atom

D

none of these

Q 52

A particle having the charge of 20 electrons on its falls through a potential difference of 100 volts.Calculate the energy acquired by it in electron volt (eV).

A

2.0 x 10^-2 eV

B

2.0 x 10^-3 eV

C

2.0 x 10^2 eV

D

2.0 x 10^3 eV

Q 53

Steady current does not change with respect to _____.

A

conductor

B

source

C

time

D

potential difference

Q 54

The activity of a certain nuclei decreases to 15 % of its original value in 10 days. Find its half life?

A

2 days

B

2. days

C

3.65 days

D

4 days

Q 55

Acceleration due to gravity near earth is

A

nonuniform

B

uniform

C

decreasing with distance

D

increasing with time

Q 56

A pipe is filled with a gas and open at one end. If the length of the pipe is 0.6 m and the speed of sound in the gas in 300 m/s. Then frequencies of the first two harmonics are:

A

125 Hz and 250 Hz

B

250 Hz and 750 Hz

C

250 Hz and 500 Hz

D

125 Hz and 375 Hz

Biology

Q 57

The alternative forms of same gene is called which of the following?

A

locus

B

gene

C

allele

D

phenotype

Q 58

Which of the following carbohydrate is a tasteless molecule?

A

Monosaccharides

B

Oligosaccharides

C

Polysaccharides

D

carbohydrates

Q 59

The only aquatic arthropods:

A

Crustaceans

B

Arachnids

C

Myriapods

D

Gastopods

Q 60

Compound Microscope was first used by

A

A.V. Leeuwenhoek

B

Pasteur

C

Janssen and Hans

D

None of these

Q 61

Which two proteins are the major components of myofibrils, allowing for muscle fiber contraction?

A

Myosin and cartilage

B

Mylode and myosin

C

Lamellae and actin

D

actin and myosin

Q 62

What percentage of protein is found in the cell membrane?

A

20-4-

B

40-50

C

60-80

D

90

Q 63

Herpes simplexes is caused by which virus?

A

adenovirus

B

Pox virus

C

influenza Virus

D

herpes virus

Q 64

In mixed inhibition, the inhibitor binds on the enzyme ?

A

allosterically

B

active site

C

doesn't bind on the enzyme

D

binds on the substrate

Q 65

Which statement is incorrect?

A

In short day plants red light prevents flowering

B

In long day plants for red light promotes flowering

C

Leaf unrolling occurs in grasses

D

Henbane is a short day plant

Q 66

As long as two species occupy different niches, there is?

A

competition

B

no competition

C

gene linkage

D

polymorphism

Q 67

Which of the following bacteria are thick rigid and spiral?

A

Coccus

B

Vibrio

C

spirillum

D

Bacillus

Q 68

The attraction between water molecules is termed as:

A

cohesion

B

tension

C

adhesion

D

imbibition

Q 69

The lighter, inner section of the brain is called:

A

white matter

B

Gray matter

C

Reflex arc

D

Medulla

Q 70

Which of the light is mainly absorbed by the plants?

A

Orange

B

Red

C

Green

D

Both A and B

Q 71

Which statement is incorrectabout ethylene production?

A

Climacteric is burst of respiratory activity in fruit ripening

B

It is associated with ethane production

C

It helps in fruit ripening

D

It helps in fruit set

Q 72

Coelom is lined by which of the following?

A

Parietal mesoderm

B

Visceral mesoderm

C

Mesoderm

D

both a and b

Q 73

If more oxygen is present, the rubisco starts:

A

Respiration

B

Photorespiration

C

Carboxylase

D

None of these

Q 74

The idea of inheritance of acquired characteristics was given by

A

Lamark

B

Darwin

C

Aristotle

D

Lyell

Q 75

Hemoglobin molecule exhibits which structural organization of proteins?

A

primary structure

B

secondary structure

C

tertiary structure

D

quaternary structure

Q 76

In which of these reflexes, a skeletal muscle contraction causes the agonist muscle to simultaneously lengthen and relax?

A

stretch reflex

B

spinal reflex

C

golgi tendon reflex

D

Crossed Extensor Reflex

Q 77

The main feature of prokaryotic cell is

A

absence of DNA

B

absence of nucleus

C

absence of cell wall

D

absence of plastids

Q 78

Nephridia are the excretory organs of members of which phylum?

A

Arthropoda

B

Cnidaria

C

Annelida

D

Mollusca

Q 79

A botanist sees that when he breeds a plants with blue flowers with a plant with red flowers, the resulting generation are plants that all have a 1 : 1 ratio of blue : red flowers. He knows that the two parents are homozygous for the trait of color. What phenomenon most likely explains the 1 : 1 ratio in the filial generation?

A

Incomplete dominance

B

Independent assortment

C

Penetrance

D

Codominance

Q 80

The internal buds are known as which of the following?

A

Spicules

B

Choanocytes

C

Gemmules

D

both a and b

Q 81

Midbrain is also known as:

A

Pons

B

Medulla

C

mesencephalon

D

None of these

Q 82

Lungs are spongy because of the presence of:

A

alveoli

B

trachea

C

diaphragm

D

bronchi

Q 83

The Induced fit model was introduced by Koshland in which of the following year?

A

1960

B

1961

C

1959

D

1966

Q 84

The muscle which moves a body part away from the midline of the body is:

A

flexor muscles

B

Extensor muscles

C

adductor muscles

D

abductor muscles

Q 85

The percentage of light absorbed by the leaf is?

A

0.2

B

0.15

C

0.05

D

0.01

Q 86

What are the total number of amino acids found in a hemoglobin molecule?

A

575

B

576

C

574

D

573

Q 87

This is non-reducing sugar

A

maltose

B

lactose

C

cellobiose

D

sucrose

Q 88

Which of the following scientists hypothesized that organisms can pass down traits acquired during their lifetimes?

A

Lamark

B

Darwin

C

Mendel

D

Linnaeus

Q 89

An entire skeletal muscle is surrounded by _____.

A

sarcolemma

B

epimysium

C

Both A and B

D

microtubules

Q 90

Why are phospholipids a major part of the lipid bilayer in plasma membranes?

A

They have a nitrogenous base in the head region.

B

They have fatty acids in the tail region.

C

They are amphipathic in nature.

D

They have a phosphate group in the head region.

Q 91

Which viruses enter the host cell as a whole?

A

plant virus

B

bacteriophages

C

animal virus

D

none

Q 92

How is the Hepatitis A virus transmitted?

A

Feaces

B

Sexual contact

C

blood

D

all

Q 93

Which statement about enzyme is incorrect?

A

Some of them consist solely of protein with no non protein part.

B

They catalyze a chemical reaction without being utilized.

C

They without their cofactor are called apoenzyme.

D

All enzymes are fibrous Proteins

Q 94

Central nervous system consists of:

A

brain and spinal cord

B

spinal nerves only

C

cerebellum and brain stem only

D

cerebrum and spinal column

Q 95

Which of the following is monohybrid cross?

A

TTYy x Ttyy

B

TT x tt

C

both a and b

D

none of these

Q 96

High fever, cold, and cough with sputum production are symptoms of

A

emphysema

B

asthma

C

pneumonia

D

bronchitis

Q 97

How many pair of ribs are present in chest wall?

A

10

B

11

C

12

D

13

Q 98

60S and 40S subunit combine to form what size particle?

A

80S

B

90S

C

100S

D

110S

Q 99

Malonic acid is an example of which type of inhibitors?

A

Irreversible inhibitor

B

Reversible inhibitor

C

Non-competitive inhibitor

D

Competitive inhibitor

Q 100

Gut in pseudocoelomates is made from which of the following?

A

Ectoderm

B

Mesoderm

C

Endoderm

D

all of these

Q 101

In seminiferous tubules repeated division of the germinal epithelium produce?

A

oogonia

B

spermatogonia

C

ovogonia

D

oogonia

Q 102

Name the first part of small intestine.

A

Duodenum

B

Jejunum

C

Ileum

D

Both A and B

Q 103

Enzymes do not change which of the following?

A

conc. Of reactants

B

conc. Of products

C

Keq

D

both a and b

Q 104

Which one of these occur in dark reactions of photosynthesis

A

formation of ATP

B

release of oxygen

C

release of hydrogen

D

synthesis of PGAL

Q 105

In icosahedral, in which the capsomers are arranged in _____ triangles

A

1

B

20

C

3

D

4

Q 106

In terrestrial conditions, the type of fertilization more common is:

A

internal

B

external

C

self

D

All A,B and C are correct

Q 107

What was the unit membrane model?

A

Plasma membrane has lipid bilayer.

B

Proteins are embedded in the lipid bilayer.

C

Plasma membrane has charged pores for transport of materials which cannot penetrate through the lipid bilayer.

D

All of the above.

Q 108

The fibrous connective tissue which attaches bone to bone is called:

A

Tendon

B

Ligament

C

reticular tissue

D

Cartilage

Q 109

Fatty acids release considerable amount of energy in oxidation and ___

A

calvin cycle

B

Kreb's cycle

C

dark reaction

D

light reactions

Q 110

Which of the following is not a true characteristic of spermatogonia?

A

They develop into primary spermatocytes through mitosis

B

They are undifferentiated

C

They are germ line cells

D

They are haploid

Q 111

Which of the following is not an evidence for evolution?

A

Fossil record

B

Common ancestor organisms

C

Vestigial structures

D

none of these

Q 112

During lytic cycle how many phages are released from infected host cell?

A

100-300

B

100-500

C

100-200

D

100-400

Q 113

The last common ancestor of humans is known to be which of the following?

A

Homo neanderthalensis

B

Lemuroidea

C

Dromaeosaurus

D

Pan troglodytes

Q 114

Which of the following comes under structural classification?

A

Synchondroses

B

sutures

C

Gomphosis

D

All of these

Chemistry

Q 115

Saponification number describes___

A

Unsaturation in fat

B

Average molecular weight of fatty acid

C

Acetyl number

D

Acid number

Q 116

Cnidaria is characterized by which of the following?

A

Tissue level of organization

B

Nematoblasts

C

Coelenteron

D

all of these

Q 117

The diencephalon comprises of:

A

Pons and medulla

B

Thalamus and limbic system

C

Pons and medulla

D

Hypothalamus and limbic system

Q 118

Which of the following is NOT a method bacteria use for genetic recombination?

A

Conjugation

B

Transformation

C

Transduction

D

Binary Fission

Q 119

Which of the modes of cellular transport requires energy?

A

active transport

B

osmosis

C

passive transport

D

diffusion

Q 120

Which group does not contain polyp?

A

Anthozoa

B

Hydrozoa

C

Scyphozoa

D

Calcarea

Q 121

Which brain part is responsible for our basic and primitive emotions?

A

Limbic system

B

Thalamus

C

hypothalamus

D

cerebrum

Q 122

Reversible inhibitors form Weak linkages with which of the following?

A

enzyme

B

reactant

C

product

D

substrate

Q 123

Conversion of ssRNA (-) to ssRNA (+) mediated by

A

DNA dependent DNA polymerase

B

. DNA dependent RNA polymerase

C

RNA dependent DNA polymerase

D

RNA dependent RNA polymerase

Q 124

The factor which decreases the oxygen saturation of hemoglobin:

A

C02

B

Temperature

C

pH of blood

D

All of these are correct

Q 125

Formula mass is considered for ___compounds instead of their molecular mass

A

metallic

B

ionic

C

covalent

D

polar covalent

Q 126

Which state of matter don't have definite volume and occupy space

A

solids

B

liquids

C

gases

D

all of these

Q 127

Iodine has greater heat of sublimation then its family members due to presence of stronger?

A

H-bonding

B

London dispersion forces

C

Dipole-dipole force

D

Chemical bonding

Q 128

If both alkyl groups attached to carbonyl in ketone are same then it is called as_____?

A

Mixed

B

Unsymmetrical

C

Symmetrical

D

None of these

Q 129

Activated complex is formed due to

A

Pressure

B

Temperature

C

Effective collision

D

Ineffective collision

Q 130

Which of the following test distinguish between 2-butanol and 1-butanol?

A

Baeyer's test

B

Lucas Test

C

Benedict's test

D

2,4-DNPH test

Q 131

Albumin is present in____?

A

Milk

B

Eggs

C

Beans

D

Muscles

Q 132

Kinetic energy of the molecule may be all of the following except?

A

Vibrational

B

Translational

C

Rotational

D

Static

Q 133

The oxidation number in [MnO4]-2

A

7

B

-7

C

6

D

-6

Q 134

Which of the following Compound is not reduced by $NaBH_4$?

A

Acetaldehyde

B

Acetone

C

Carboxylic acid

D

Alkene

Q 135

Equilibrium constant has

A

Units

B

No Units

C

Both A and B

D

A negative value

Q 136

Chemical reactions involve change in heat energy and the study is called

A

Electrochemistry

B

Biochemistry

C

Thermochemistry

D

Analytical Chemistry

Q 137

Energy is __when an electron is added to the isolated gaseous atom

A

increased

B

released

C

decreased

D

absorbed

Q 138

OH-is added when reaction is in

A

Acidic Medium

B

Basic Medium

C

Neutral solution

D

All of these

Q 139

The number of charges present on a cation depends on number of electron _by the atom

A

gain

B

lost

C

accept

D

produced

Q 140

Which of the following is an example of Tricarboxylic acids ?

A

Maleic acid

B

Citric acid

C

Butyric acid

D

None of these

Q 141

Heavy water contain isotope of Hydrogen

A

protium

B

deuterium

C

tritium

D

both A & C

Q 142

Which of the following reaction of carbonyl system is used to distinguish between carbonyl system from alcohols?

A

Cyanohydrin formation

B

Bisulfite adduct formation

C

Haloform formation

D

None of these

Q 143

The reaction in Galvanic Cell is

A

Spontaneous

B

nonspontaneous

C

Irreversible

D

endothermic

Q 144

The intermolecular forces brings the molecules close together and give them particular

A

chemical properties

B

physical properties

C

composition

D

nature

Q 145

in solid iodine, I-I bond distance is

A

77.1 pm

B

271.5 pm

C

11 pm

D

166.7 pm

Q 146

An example of non- polar molecular crystal is

A

Ice

B

Iodine

C

Sugar

D

Salt

Q 147

The substances which reduces the activity of enzyme action are called as_____?

A

Reducers

B

Activators

C

Promotors

D

Inhibitors

Q 148

when water freezes to ice, it occupies ___ more space

A

0.9 g/cm^3

B

0.1 g/cm^3

C

0.12 g/cm^3

D

0.13 g/cm^3

Q 149

When ethyl alcohol is oxidized with potassium dichromate, which product is formed?

A

Acetic acid

B

Butanoic acid

C

Benzoic acid

D

No reaction takes place

Q 150

Graphite has a structural similarity with

A

B2H6

B

B4C

C

B

D

BN

Q 151

Ethene react with Oxygen in the presence of Silver oxide to produce_____?

A

Ethylene oxide

B

Ethylene epoxide

C

Methylene oxide

D

Both a and b

Q 152

The imaginary surface separating system and surroundings is called

A

Buffer

B

Transition Zone

C

Boundary

D

Intermediate state

Q 153

Carboxylic acids having long aliphatic chain are called as_____?

A

Long carboxylic acids

B

Higher carboxylic acids

C

Fatty acids

D

Glycerols

Q 154

All the halogens are _____diatomic molecules

A

polar

B

non polar

C

reactive

D

non reactive

Q 155

Both Hydrogen and Halogens form

A

Ionic compounds mostly

B

Unipositive ion

C

Diatomic molecule

D

Stable bond with water

Q 156

Which one of the following is an example of transition element?

A

Na

B

Co

C

Ba

D

Ra

Q 157

Which inert gas is mixed with oxygen gas by deep sea divers to adjust the partial pressure of oxygen gas

A

Neon

B

Helium

C

Argon

D

Krypton

Q 158

The value of energy obtained for the electron in the nth orbit of hydrogen atom is in

A

Joules

B

Joules/atom

C

Kilojoules

D

Kilojoules/atom

Q 159

Alkanes mostly give_____?

A

Addition reactions

B

Elimination reactions

C

Reduction reactions

D

Substitution reactions

Q 160

Instantaneous dipole is produced when _____?

A

Two polar molecules comes closer

B

Two non polar molecules comes closer

C

Polar and nonpolar molecule comes closer

D

All of these

Q 161

The order of reaction for which half-life is inversely proportional to the concentration of reaction is

A

Second Order

B

Zero Order

C

First Order

D

Third Order

Q 162

Formaldehyde polymerizes to form ?

A

Bakelite

B

Paraldehyde

C

Metaformaldehyde

D

All of these

Q 163

In the reaction; Zn + $CuSO_4$ → $ZnSO_4$ + Cu, Which of the following change shoes that redox reaction takes place at once ?

A

Blue color appears

B

Blue color disappears

C

Temperature of the surrounding decreases

D

None of these

Q 164

The method which is used on industrial scale for production of alcohol from alkene ?

A

Hydrohalogenation of alkenes

B

Dehydration of alkenes

C

Hydration of alkenes

D

Hydroxylation of alkenes

Q 165

When an electron changes its orbit from outer to inner level, energy is

A

no change

B

absorbed

C

released

D

remains constant

Q 166

I elements have high ionization energy value then it means they belongs to

A

metals

B

non metals

C

metalloid

D

noble gases

Q 167

The isomerism in which compounds have different number of carbon atoms on both sides of the functional group is called as?

A

Tautomerism

B

Metamerism

C

Geometrical Isomerism

D

Position isomerism

Q 168

Alkali and alkaline earth metals impart colours when heated over the burner, it is due to

A

Smaller electronegativity of alkali metals

B

The smaller ionic radius of these metals

C

De-excitation of electrons from higher energy levels to low energy level

D

Excitation of electrons from low energy levels to higher energy levels

Q 169

In electrolysis of $NaNO_3$, Nitrate ions are not discharged but

A

Hydroxide Ions are discharged

B

Hydrogen Ions are Discharged

C

Sodium Ions are Discharged

D

None of these

Q 170

The _study of composition of pure substance in 17th century clearly shows that few elements are components of many substances

A

qualitative

B

quantitative

C

both A & B

D

extensive

Q 171

Enzymes that catalyze the addition of ammonia, water or carbon dioxide to double bond or their removal are called_____.

A

Lyases

B

Hydrolases

C

Ligases

D

Transferases

Q 172

Vital force theory proposed by

A

Wohler

B

Berzelius

C

lanthanides and actinides resemble in?

D

Lyll

Q 173

What is the correct equation to find Heat of neutralization using glass calorimeter ?

A

q=m s ΔT

B

E=q m ΔT

C

E=m s ΔT

D

q=m c ΔT

Q 174

At equilibrium if the concentration of product is increased reaction will proceed to

A

Forward Direction

B

Backward Direction

C

Remain Undisturbed

D

None of these

Q 175

The total number of transition elements are?

A

48

B

32

C

58

D

28

Q 176

The type of hybridization in diamond is

A

spherical

B

sp2

C

sp3

D

sp-sp

Q 177

line spectrum of sodium contains ____ colored lines separated by a definite distance

A

one yellow

B

two yellow

C

two brown

D

two golden

Q 178

Which of the following shows H-bonding?

A

CH_3CH_2OH

B

CH_3-O-CH_3

C

CH_3CH_2Cl

D

All of these

Q 179

The least value of Van der Waals constant is of

A

H_2

B

N_2

C

CO_2

D

Cl

Q 180

What is the correct order of reactivity of alkyl halides ?

A

R-Cl>R-Br>R-F>R-I

B

R-I>R-Br>R-Cl>R-F

C

R-I>R-Cl>R-Br>R-F

D

None of these

English

Q 181

I found the (A)/ two first chapters (B) /of the book (C)/ particularly interesting. (D)

A

I found the

B

two first chapters

C

of the book

D

particularly interesting.

Q 182

Thank you. You've been very _____ through this time.

A

ideal

B

cruel

C

kind

D

glad

Q 183

The thief has ____ all my belongings.

A

robbed

B

taken

C

stolen

D

taken away

Q 184

Choose the correct sentence.

A

Belgium and france fought over who would host the olympic games?

B

Belgium and France fought over who would host the Olympic Games.

C

Belgium and France: fought over who would host the Olympic Games!

D

Belgium and France fought over who would host the olympic games.

Q 185

Do you wear _____ uniform to school?

A

a

B

an

C

the

D

no article

Q 186

He _____ books these days.

A

sells

B

sell

C

has sold

D

sold

Q 187

Can you pass me _____ salt?

A

a

B

an

C

the

D

no article

Q 188

Both of my roommates _____ decided to live in the dorms.

A

has

B

have

C

had

D

will have left

Q 189

The old man _____ (yell) at the boys who teased him.

A

yell

B

yelled

C

yelling

D

yells

Q 190

Apart from cows, buffaloes are also _____ for milk production.

A

reared

B

grown

C

developed

D

produced

Q 191

Identify the tense used in the given sentence. “Everyone shall be reaching by tomorrow.”

A

Present

B

Past

C

Future

D

None

Q 192

Would you mind _____ (open) the window please, it's hot here.

A

open

B

opens

C

opened

D

opening

Q 193

he is

A

hes

B

hes'

C

he's

D

heis

Q 194

Choose the correct spelling of the word

A

disappeared

B

disapeared

C

disapared

D

disapearred

Q 195

We are trying (A)/ to locate the (B) /historical city for(C) /the past two years. (D)

A

We were trying

B

to locate the

C

historical city for

D

the past two years.

Q 196

Each and every member _____ to vote.

A

has

B

have

C

having

D

are

Q 197

Choose the correct sentence.

A

Qamar is going to college in September, he plans to study Sociology.

B

Qamar is going to college in september, he plans to study sociology.

C

Qamar is going to college in September he plans to study Sociology.

D

Qamar is going to college in september he plans to study Sociology?

Q 198

A pack of lions ____ approaching the camp.

A

will

B

were

C

are

D

was

Q 199

I _____ (write)a new book. I _____ (hope) to finish it by the end of this month.

A

write, hoping

B

am writing, hope

C

wrote, am hopeful

D

writing, hope

Q 200

This summer our vacation should be both exciting and restful.

A

complex

B

simple

C

compound

D

compound-complex

Logical Reasoning

Q 201

Fact 1 Ayesha said Hamza and I both have cats Fact 2 Hamza said I don’t have a cat Fact 3 Ayesha always tells the truth but hamza sometimes lies If the above three statements are facts than which of the following statement will also be a fact I. Ayesha has a cat II. Hamza has a cat III. Hamza is lying IV. All the statements are the facts

A

Only 1

B

Only II

C

Only III

D

Statement 4

Q 202

Statements (I) The childrens are being aggressive today. (II) There is no specific limit for childrens to use mobile screens

A

Statement 1 is the cause and 2 is the effect

B

both of the statements 1 and 2 are independent.

C

Statement 2 is the cause and 1 is the effect

D

Both statements 1 and 2 are effects of some common cause

Q 203

Which one of the following have four sides

A

triangle

B

square

C

circle

D

right triangle

Q 204

Complete the series BXM, EUP, HRS, _.

A

XZY

B

RPN

C

VTU

D

KOV

Q 205

Tea is to mug as the jam is to

A

Plate

B

Bread

C

Spoon

D

Water

Q 206

Statement The Management of School M has decided to give free breakfast from next academic year to all the students in its primary section through its canteen even though they will not get any government grant. Courses of Action (I) The school will have to admit many poor students who will seek admission for the next academic year. (II) The canteen facilities and utensils have to be checked and new purchases to be made to equip it properly. (III) Funds will have to be raised to support the scheme for years to come.

A

Only II and III follows

B

Only III and I follow

C

Only I and II follow

D

Only I follows

Q 207

If a computer has more than one processor then it is known as ?

A

Uni-process

B

Multiprocessor

C

Multi-threaded

D

Both A and B

Q 208

Renown , Reputation, Abeyance, ?

A

Misery

B

Humble

C

Discontinuation

D

Best

Q 209

The SI unit of charge is ?

A

Ampere

B

B. Coulomb

C

C. Ohm

D

Both A and B

Q 210

Statement Should coal engines be replaced by electric engines in trains? Arguments (I) Yes. Coal engines cause a lot of pollution. (II) Yes. Electric engines are good on performance, easy to operate, and low on maintenance. (III) No. Pakistan does not produce enough electricity to fulfill its domestic needs also."

A

All are strong

B

Only I and II are strong

C

Only II and III are strong

D

Only I and III are strong

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