

Paid Practice Test Bundle 2 Paper 3

Physics

Q 1

SI unit of acceleration is:...

A

m

B

m²

C

m/s

D

m/s²

Q 2

Radar system is an application of:...

A

Doppler's effect

B

mechanical effect

C

electric effect

D

magnetic effect

Q 3

Which ray shows comparable penetrating power to x rays

A

alpha

B

beta

C

gamma

D

white

Q 4

What is 1 radian in degrees approximately

A

57.3 degrees

B

360 degrees

	C
π degrees	
	D
π^2 degrees	
Q 5	
When $N_s > N_p$ then transformer is	
	A
step up	
	B
step down	
	C
primary	
	D
secondary	
Q 6	
An electrical instrument which is used to measure potential difference between two points is called	
	A
Barometer	
	B
Manometer	
	C
Galvanometer	
	D
Voltmeter	
Q 7	
_____ relationship exists between current and voltage in terms of Ohm's law	
	A
Non linear	
	B
Varying	
	C
Linear	
	D
None of them	
Q 8	
When total work done on a particle is positive then	
	A
KE remain constant	
	B
momentum increases	
	C
KE decreases	

D

all of these

Q 9

Phase difference between two particles of a medium lying between two consecutive nodes is:

A

Zero

B

$\pi/2$

C

$\pi/4$

D

π

Q 10

β^- decay means emission of electron from

A

radioactive nucleus

B

innermost electron orbit

C

a stable nucleus

D

outer most electron orbit

Q 11

Which of the following statement is not true about heat engine?

A

All real engines are less efficient than Carnot engine

B

All real engines are less efficient due to friction and heat losses

C

efficiency of Carnot engine working between same two temperatures, depends on the nature of working substance

D

the larger the temperature difference of two reservoirs, the greater is the efficiency

Q 12

Which of the following substances cannot be emitted by radioactive substances during their decay?

A

protons

B

neutrinos

C

helium nuclei

D

electrons

Q 13

The resistance and length of wire are

A

inversely related

B

directly related

C

not related

D

inversely proportional

Q 14

When work done by force of gravity is negative (only gravity is acting)

A

KE decreases

B

PE increases

C

both (A) and (B)

D

none of these

Q 15

The photon is the particle , which has :

A

Infinite rest mass

B

Rest mass but no charge

C

No rest mass & no charge

D

A & C are correct

Q 16

The number of diodes in bridge rectifier is

A

4

B

3

	C
2	
	D
5	

Q 17	
What will be the decay constant of 1 Curie sample of radioactive substance of mass 214, its half life is 26.8 min?	
	A
4.31×10^{-4}	
	B
4.31×10^{-5}	
	C
4.31×10^5	
	D
0.431	

Q 18	
The four bulbs of 40 W each are connected in series with a battery across them. Which of the following statement is true?	
	A
The current through each bulb is same	
	B
The voltage across each bulb is not same	
	C
The power dissipation in each bulb is not same	
	D
None of the above	

Q 19	
The star which is very hot will emit _____ color	
	A
Blue	
	B
red	
	C
orange	
	D
green	

Q 20	
Atomic spectra is a _____ spectra	
	A
continuous	

	B
discrete	
	C
both a) and b)	
	D
None of these	
Q 21	
If a photon is absorbed by a nucleus the energy of nucleus	
	A
remain same	
	B
increase slightly	
	C
decrease slightly	
	D
it will pass the nucleus	
Q 22	
Do magnetic flux lines intersect?	
	A
Yes	
	B
No	
	C
depends on strength of field	
	D
cannot be determined	
Q 23	
A cycle tyre bursts suddenly is an example of	
	A
Isothermal process	
	B
Isochoric process	
	C
adiabatic process	
	D
Isobaric process	
Q 24	
Atom is neutral because it has equal number of	
	A
charge particles	
	B
uncharged particles	

	C
neutrons	
	D
all of these	

	Q 25
The coulomb constant is defined as:...	
	A
$1 / (4\pi\epsilon)$	
	B
$4 / (\pi\epsilon)$	
	C
$4\pi\epsilon$	
	D
$\pi / (4\epsilon)$	

	Q 26
When a man walks on a surface horizontally with constant velocity, work done by friction is zero	
	A
contact force is zero	
	B
gravity is zero	
	C
all of these	
	D

	Q 27
What is the measure in radians of the angle $A = 330^\circ$?	
	A
$11\pi/3$	
	B
$7\pi/4$	
	C
$7\pi/6$	
	D
$11\pi/6$	

	Q 28
The direction of induced current is always so as to oppose the change which causes the current is	
Faraday's Law	
	A
	B

Lenz's Law

C

Ohm's Law

D

Kirchoff's Law

Q 29

The terminal potential difference of a battery is greater than its emf when

A

The internal resistance of a battery is infinite

B

The internal resistance of a battery is zero

C

The battery is charged

D

The battery is discharged

Q 30

___ work on arbitrary system means a transfer of energy to the system

A

Positive

B

negative

C

Can be positive or negative

D

None of the above

Q 31

Determine the linear velocity of a point rotating at an angular velocity of 12π radians per second at a distance of 8 centimeters from the center of the rotating object.

A

31.6 cm/s

B

301.6 cm/s

C

30.6 cm/s

D

3016 cm/s

Q 32

Find the Lorentz force of a charge 2.5C having an electric field of 5 units and magnetic field of 7.25 units with a velocity 1.5m/s.

A

39.68

	B
68.93	
	C
89.39	
	D
63.98	

Q 33	
An electric filament bulb can be worked from	
	A
D.C. supply only	
	B
A.C. supply only	
	C
Battery supply only	
	D
All above	

Q 34	
The linear acceleration of a body moving in a circular path is:	
	A
negative	
	B
positive	
	C
constant	
	D
zero	

Q 35	
The maximum efficiency of full wave rectifier is	
	A
80.60%	
	B
40.60%	
	C
70%	
	D
50%	

Q 36	
Michelson determined velocity of light by	
	A
studying rotation of moon and sun	
	B

using interferometer

C

using a rotating plane mirror

D

using an octagonal rotating mirror D

Q 37

Centripetal acceleration always acts _____ the center

A

away

B

towards

C

normally

D

tangentially

Q 38

A gas expands from V_1 to V_2 . The amount of work done is greatest in

A

Adiabatic

B

Isobaric

C

Isochoric

D

Isothermal

Q 39

A wave passes through a medium, each particle of the medium performs 100 complete vibrations in 5 seconds. What is the frequency of the wave:

A

2 Hz

B

20 Hz

C

4 Hz

D

40 Hz

Q 40

x-rays are produced in _____ tube

A

Ar filled

B

He filled

C

H filled

D

Vacuum

Q 41

If a squash ball comes back to its starting point after bouncing off the wall several times, then:....

A

its total displacement is zero but its average velocity is non-zero

B

its total displacement is non-zero but its average velocity is zero

C

both its total displacement and its average velocity is non-zero

D

its total displacement is zero and so also is its average velocity

Q 42

In a standing waves, the distance between two consecutive nodes is:

A

equal to one wavelength

B

equal to two wavelength

C

equal to half of wavelength

D

equal to quarter of wavelength

Q 43

To get a peak load voltage of 40V out of a bridge rectifier, what should be the approximate rms value of secondary voltage?

A

0V

B

14.4V

C

28.3V

D

56.6 V

Q 44

The white laser cannot be produced because

A

its is not coherent

B

it has low energy

C

it diffracts easily

D

all of these

Q 45

4000 Coulomb charges were passing from the wire for about 12 seconds. Estimate the current during this process?

A

333.3 ampere

B

333.33 volts

C

666.67 ampere

D

None of them

Q 46

For a car which applies brakes from 10 m/s to stop the car in 10 sec its acceleration is

A

1 m/s²

B

2m/s²

C

-1 m/s²

D

-2 m/s²

Q 47

During an adiabatic process pressure of gas is found to be proportional to the cube of its temperature. The ratio of C_p/C_v is

A

2

B

44319

C

44257

D

44289

Q 48

An object is displaced from position vector $r_1 = (2i + 3j)m$ to $r_2 = (4j + 6k)m$ under a force $F = (3x^2 i + 2y j) N$. Find the work done by this force

A

55 J

B

83 J

C

0

D

-83 J

Q 49

What is the wavelength of the wave if the phase angle between two points of the medium is $3\pi/4$ and they are separated through a distance of 3 cm?

A

8 cm

B

9 cm

C

1 cm

D

12 cm

Q 50

A reversible carnot engine converts 1/6th of heat into input work. When the temperature of sink is reduced by 62 degree C then efficiency is doubled then temperature of source and sink is

A

80 C ,37 C

B

99 C , 30C

C

99C ,25C

D

99C , 37C

Q 51

A wire has a resistance of 5.5Ω at 19°C and 21.5Ω at 200°C . Find the temperature coefficient of resistivity(α) of the material.

A

0.016 per degree celsius

B

32 per degree celsius

C

0.018 per degree celsius

D

0.00106 per degree celsius

Q 52

Which color of heat shows highest temperature

A

red

B

blue

C

black

D

orange

Q 53

The energy stored in a parallel plate capacitor is 24 J. What is the potential difference between the plates if the capacitance of the capacitor is 3 μF ?

A

4 kV

B

16 kV

C

54 kV

D

8 kV

Q 54

A beam of ion with velocity 2×10^5 m/s enters normally into a magnetic field of 0.04 T. The specific charge of ion is 5×10^7 C/kg. Radius of circle is

A

0.1 m

B

0.16 m

C

0.2 m

D

0.25 m

Q 55

The unstable atom means

A

electrons are increasing

B

protons are increasing

C

neutrons are increasing

D

any of these

Q 56

Two tuning forks produces N beats. If one of these tuning forks has the frequency f, then the frequency of the other would be:

A

$N - f$

B

N / f

C

$N * f$

D

N + f

Biology

Q 57

The thalamus and the hypothalamus are located in which region of the brain?

A

brain stem

B

cerebrum

C

cerebellum

D

Diencephalon

Q 58

An enzyme without its cofactor is called _____ .

A

Coenzyme

B

Apoenzyme

C

Holoenzyme

D

Proenzyme

Q 59

Enzyme-catalyzed modifications are?

A

reversible

B

irreversible

C

both a and b

D

none of these

Q 60

The flagella originates from which part of the cell?

A

basal body

B

cell membrane

cell wall	C
capsule	D
Its protein in nature	Q 61
fats /cholesterol	A
ATP	B
not known without screening test	C
Ligase	D
What are the Phages that show lysogenic cycle called?	Q 62
Lytic phages	A
virulent phages	B
Temperate phages	C
None of these	D
Evaporation of water from aerial parts of plants is called:	Q 63
ascent of sap	A
deplasmolysis	B
deplasmolysis	C
none of these	D
Resolution of human eye is:	Q 64
2.0 μm	A
2.0 mm	B
1.0 μm	C

D

1.0 mm

Q 65

Bottlenecks increases the affect of which of the following?

A

genetic linkage

B

genetic expression

C

genetic diversity

D

gene pool

Q 66

Rate of photosynthesis does not depend upon:

A

Quality of light

B

Intensity of Light

C

Duration of Light

D

Temperature

Q 67

The name animal is derived from what word?

A

Aname

B

Anemia

C

Anima

D

none of these

Q 68

Which theory tells about adaptation

A

Darwin's natural selection

B

Lamarck's theory

C

Hardy

D

Weinberg's principle

Q 69

What is the specialized organelle in muscle cells that serves to sequester calcium?

A

T-tubule

B

Sarcolemma

C

Sarcomere

D

Sarcoplasmic reticulum

Q 70

The first layer of cell wall which is formed is called?

A

primary wall

B

secondary wall

C

middle lamella

D

all of these

Q 71

Carapace is present in which class of arthropoda?

A

Arachnids

B

Insects

C

Crustaceans

D

All of them

Q 72

Genes evolve due to which of the following?

A

artificial selection

B

random selection

C

gene selection

D

natural selection

Q 73

The graph that shows relative effectiveness of different wavelengths in photosynthesis is?

A

actin spectrum

B

action spectrum

C

absorption spectrum

D

emission spectrum

Q 74

If the concentration of enzyme is kept constant and amount of substrate is increased a point is reached where increase in substrates concentration does not affect the reaction rate because of?

A

Enzymes get denatured at higher substrate conc.

B

Rate of reaction is indirectly proportional to substrate concentration at this point.

C

All the active sites on enzyme molecule are occupied

D

all of these

Q 75

Trachea is also termed as:

A

voice box

B

epiglottis

C

bronchi

D

windpipe

Q 76

Such inheritance in which traits vary quantitatively is?

A

Continuously varying trait

B

Incomplete dominance

C

Test cross

D

Polygenic inheritance

Q 77

How many bones do humans have in the vertebral collumn?

A

B

25

C

33

D

34

Q 78

Which of the following is an example of a marine annelid?

A

Neries

B

Stylaria

C

Earthworm

D

Leech

Q 79

If we consider Hardy-Weinberg law, then following is incorrect in its sense?

A

Mutations cause changes in genetic frequency

B

Migration changes allele frequency

C

There should not be selection

D

Non-random mating will reduce chances of evolution

Q 80

The element in basic structure of proteins which differentiate them from carbohydrates is?

A

C

B

H

C

N

D

none of these

Q 81

Which among the following defines GPI anchored proteins?

A

Integral proteins of the plasma membrane

B

Proteins that bind to ion-gated channels in the plasma membrane

C

Proteins which randomly bind to lipids of the plasma membrane

D

Peripheral proteins of the plasma membrane

Q 82

The Herpes virus is responsible for which of the following types of Herpes?

A

simplex

B

duplex

C

triplex

D

quadruplex

Q 83

Which one of the following does not occur at puberty in girls?

A

voice deepens

B

pubic hair grow

C

hips broaden

D

menstrual cycle starts

Q 84

While bound to the active site, the substrate is converted into which of the following?

A

complex

B

substrate of high energy

C

product of reaction

D

both a and b

Q 85

Some auditory reflexes and posture are controlled by which region of the brain?

A

forebrain

B

spinal cord

C

hindbrain

D

midbrain

Q 86

Which of the following is a mesh of interconnected membranes involved in protein synthesis and transport?

A

ER

B

cytoskeleton

C

golgi apparatus

D

all of these

Q 87

Where in the female reproductive system thousands of eggs are found?

A

ovaries

B

vagina

C

fallopian tube

D

uterus

Q 88

Which of the following are assembled and disassembled during the cell cycle?

A

microtubules

B

microfilaments

C

all of these

D

none of these

Q 89

Which of the following would not be found in a prokaryotic cell?

A

mitochondria

B

RNA

C

Ribosomes

D

plasma membrane

Q 90

The number of linkage groups in humans is?

A

1/21

B

23

C

1/23

D

1/24

Q 91

The spinal cord is the continuation of which of the following?

A

cerebellum

B

pons

C

medulla oblongata

D

brain stem

Q 92

The respiratory chain electron carrier that directly binds oxygen is ____

A

Cytochrome a

B

cytochrome a₃

C

cytochrome c₁

D

cytochrome b

Q 93

An egg is fertilized in laboratory and implanted in uterus for development. This is called:

A

test tube baby

B

Cloning

C

In vivo fertilization

D

Both A and B

Q 94

When was the Tobacco mosaic virus successfully crystallized?

A

1935

B

1930

C

1932

D

1920

Q 95

The organs of locomotion in annelids are which of the following?

A

Muscles

B

Hydrostatic skeleton

C

Parapodia

D

Setae

Q 96

Thymus, spleen, tonsils and adenoids produce:

A

erythrocytes

B

lymphocytes

C

phagocytes

D

thrombocytes

Q 97

How many molecules of Carbon dioxide enter the calvin cycle to produce one molecule of carbohydrate?

A

2

B

3

C

4

D

1

Q 98

The contractile protein of skeletal muscle involving ATPase activity is:

A

actin

B

myosin

C

troponin

D

tropomyosin

Q 99

The fibrous connective tissue which attaches muscle to bone is called:

A

Tendon

B

Ligament

C

reticular tissue

D

Cartilage

Q 100

In a dihybrid cross, what fraction of offspring will display both recessive phenotypes?

A

1/8

B

1/4

C

1/32

D

1/16

Q 101

The outer portion of spinal cord is composed of:

A

White matter

B

Gray matter

C

Cell bodies

D

non-myelinated nerve fibres

Q 102

The epiglottis, a flap of tissues covers the

A

pharynx

B

larynx

glottis	C
nasal cavity	D
Q 103	
Which cell is the longest in the human body?	A
muscle cells	B
nerve cells	C
bone cells	D
gland cells	D
Q 104	
Fertilization of ovum occurs during which of the following?	A
in uterus	B
in ovary	C
In distal part of oviduct	D
In proximal part of oviduct	D
Q 105	
Purple non-sulphur bacteria is an example of which of the following?	A
Heterotrophic bacteria	B
Saprotrophic bacteria	C
Chemosynthetic bacteria	D
Photosynthetic bacteria	D
Q 106	
Which of the following fishes contained lungs?	A
lamprey	B
perch	C
plaice	C

D

dipnoi

Q 107

In the evolutionary sense, which organism has the highest fitness?

A

A sterile mule that can pull over 800 pounds

B

A childless human male who lives to be over one hundred years old

C

A dog who cannot give birth due to a hip abnormality, but is healthy in all other respects

D

A prairie dog that, though smaller than the average member of her species, has twice as many healthy young in each litter

Q 108

Which of the following cell structure contains the highest concentration of RNA?

A

centriole

B

nucleus

C

nucleolus

D

mitochondria

Q 109

The function of coelom is best characterized as?

A

To increase the size of the animals

B

To help in the functioning of the reproductive system

C

To provide space for the development of organs and systems

D

none of these

Q 110

The oldest mineral discovered so far is which of the following, which dates back to 4.4 billion years.

A

iron

B

zircon

C

diamond

D

cadmium

Q 111

Viruses without nuclear envelope is called as?

A

icosahedral virus

B

naked virus

C

enveloped virus

D

bilayer virus

Q 112

Who proposed pressure flow theory?

A

Dixon

B

Mohl

C

Sanger

D

Munch

Q 113

Which of the following substance is most favorable to form structural component of biological membranes?

A

Hydrophilic Carbohydrates

B

Hydrophobic fats

C

both a and b

D

none of these

Q 114

Omnis cellula e cellula is hypothesized by?

A

schleiden

B

lorenz oken

C

louis pasteur

D

rudolph virchow

Q 115

Which of the following statement about neuron is incorrect?

A

They not only conduct impulses but also generate them

B

They are not the only cellular component of nervous system

C

They may show limited regenerative capabilities

D

Like all the living cell, when they mature and divide to form similar cells

Q 116

The optimum pH for enzyme arginase is?

A

9

B

9.3

C

9.7

D

10

Q 117

Out of the 6 molecules of G3P, how many molecules are used to make glucose?

A

1

B

2

C

3

D

6

Q 118

Chitinous Setae are the locomotory organs of annelids which are present on?

A

cell wall

B

prostomium

C

nucleolus

D

parapodia

Q 119

Which of the following statements about the systemic affect of bacterial cells are true?

A

Gram-negative bacteria are more likely to cause systemic effects because their lack of an outer membrane means they can more quickly infect host cells.

B

Gram-negative bacteria are more likely to cause systemic effects because their outer membrane protects them from antibiotics

C

both a and b

D

none of these

Q 120

The epididymis, vas deferens, and urethra are a series of ducts found in which body system?

A

Endocrine

B

Lymphatic

C

Digestive

D

Male reproductive

Q 121

The Urey-Miller experiment determined which of the following results?

A

DNA replicates by semiconservative replication

B

Cyanobacteria were responsible for the oxygenation of the atmosphere

C

The early atmosphere was composed of ammonia and methane

D

Organic molecules can arise from inorganic precursors

Q 122

Which hormone is involved in opening and closing of stomata?

A

citric acid

B

Oxaloacetic acid

C

abscisic acid

D

None of these

Q 123

What types of viruses contain the enzyme lysozyme to aid in their infection?

A

bacteriophage

B

animal viruses

C

plant viruses

D

fungal viruses

Q 124

Central nervous system is present in:

A

Asymmetrical animals

B

Bilaterally symmetrical animals

C

Radially symmetrical animals

D

B and C both are correct

Chemistry

Q 125

Which of the following are product of following reaction in the presence of sunlight? $\text{CH}_4 + \text{Cl}_2 \rightarrow ?$

A

CHCl_3

B

CCl_4

C

$\text{CH}_2\text{Cl}_2, \text{CH}_3\text{Cl}$

D

All of these

Q 126

The product in SN_2 reaction is formed with _____?

A

Inversion in configuration

B

Retention in configuration

C

50% retention in configuration

D

50% inversion in configuration

Q 127

Both ketones and aldehyde are present in _____?

A

Sugars

Menthone	B
Camphor	C
Formamint	D
Q 128	
Change in concentration of reactants and products can only be found by	A
Physical methods	B
Chemical Methods	C
Physical and Chemical methods	D
None of These	
Q 129	
Salt Bridge is used for the purpose of	A
Producing Electrons	B
Circuit Completion	C
Increasing speed of electrons	D
All of these	
Q 130	
Compounds having C and H atoms and their derivatives are called as_____?	A
Inorganic compounds	B
Organic compounds	C
Biochemical compounds	D
Carbohydrates	
Q 131	
$[\text{Fe}(\text{CN})_6]^{4-}$ is called as_____?	A
Complex compound	B
Transition element ion	

C

Anion

D

Complex anion

Q 132

The specific site at which substrate is attached on the enzyme and converted into product is called as___?

A

Reaction site

B

Active site

C

Binding site

D

None of these

Q 132

The specific site at which substrate is attached on the enzyme and converted into product is called as___?

A

Reaction site

B

Active site

C

Binding site

D

None of these

Q 133

Which of the statement is incorrect for XeO_4 ?

A

four $p\pi-d\pi$ bonds are present

B

four $sp^3 - p \sigma$ bonds are present

C

It has a tetrahedral shape

D

It has a square planar shape

Q 134

Following is termed as hydroxyl derivative of alkanes

A

carboxylic

B

Aldehydes

C

Phenols and ethers

D

Alcohols

Q 135

The hydrogen gas bubbling into one molar solution of HCl has a pressure of

A

2atm

B

780 mmHg

C

19 Psi

D

1 atm

Q 136

The half-life of Uranium is

A

700 Million years

B

706 Million years

C

89 days

D

710 million year

Q 137

Hydrogenation of alkenes takes place in the presence of ____?

A

Nickel

B

Gold

C

Palladium

D

Raney Nickel

Q 138

The specie which is in search of Positive charge is called as _____?

A

Electrophile

B

Nucleophilic

C

Nucleophile

D

Cation

Q 139

Which of the carboxylic acid is used in medicine as local irritant;

A

formic acid

B

acetic acid

C

benzoic acid

D

amino acid

Q 140

Which of the following causes the inactivation of enzymes _____?

A

Concentration of substrate

B

Optimum temperature

C

Beta radiation

D

Optimum pH

Q 141

In oxidation number method of Balancing the first step is to write

A

Oxidation number on Reactants

B

Oxidation number on products

C

Oxidation number for both reactants & Products

D

None of these

Q 142

Proteoses enzyme & peptones belong to which type of protein?

A

Simple protein

B

Derived proteins

C

Conjugated proteins

D

All of these

Q 143

Which of the following is formed when phenol reacts with acetyl chloride?

A

Alcohol

B

Carboxylic acids

C

Ester

D

Amines

Q 144

On reacting with metals carboxylic acid produces which of the following products ?

A

Salt + water

B

Salt

C

Salt + H₂ gas

D

Salt + Alcohol

Q 145

Aldehyde consist of which functional group?

A

Aldehyde

B

Alcohol

C

Both a and b

D

Ketone

Q 146

Which of the following has least polarizability ?

A

NH₃

B

H₂O

C

HF

D

CH₄

Q 147

Which one of the following is not an example of state function?

A

Temperature (T)

B

Volume (V)

C

Enthalpy (E)

D

Heat (q)

Q 148

Fractions of Crude petroleum can be obtained by using_____?

A

Destructive distillation

B

Fractional distillation

C

Vacuum distillation

D

Distillation

Q 149

The maximum probability of finding an electron is at distance of

A

0.53mm

B

0.53nm

C

0.153nm

D

1.53mm

Q 150

The solubility of alcohols is due to

A

dipole moment

B

covalent bonds

C

hydrogen bonding

D

electronegativity

Q 151

What pressure of Oxygen is maintained inside the bomb calorimeter?

A

200 atm

B

25 atm

C

20 atm

D

75 atm

Q 152

During combustion analysis $MgClO_4$ is used to absorb _____

A

oxygen gas

B

hydrogen gas

C

water vapors

D

alcohol

Q 153

Acetaldehyde oxidation will lead to formation of

A

Acetic acid

B

Butanoic acid

C

Propanoic acid

D

Ester

Q 154

Which of the following is best method to prepare alkyl halides from alcohols?

A

Reaction of alcohol with HX

B

Reaction of alcohol with SOCl_2

C

Reaction of Alcohol with PCl_3

D

Reaction of Alcohol with PCl_5

Q 155

Which is un1 for phenols

A

a colourless crystalline solid

B

more reactive to electrophilic attack

C

form pink solution at room temperature

D

dissolves readily in acids

Q 156

The correct electronic configuration of Cr is

A

$[\text{Ar}]4s^23d^4$

B

$[\text{Ar}] 4s^23d^4$

C

$[\text{Ar}]4s 3d^5$

D

[Ar]4s13d5

Q 157

d-block elements are present _____ in the periodic table ?

A

Right of the periodic table

B

Left of the periodic table

C

Bottom of periodic table

D

Between s and p block elements

Q 158

The elements of which group show abnormally very low values of electron affinity in every period of periodic table

A

group 2A

B

group 5A

C

both A & B

D

none of these

Q 159

When does it mean, when a reaction is exothermic?

A

Energy content of product is more

B

Energy content of reactant is less

C

Heat is transferred from the system to surrounding

D

Heat is transferred from the Surrounding to the system

Q 160

In which direction Cathode rays deflected in the presence of magnetic field?

A

Moves upward

B

Moves downward

C

Move randomly

D

Moves in straight line

Q 161

More energy to remove an electron from _____ ?

A

Half filled subshell

B

Completely Filled Subshell

C

Partially filled subshell

D

Both a and b

Q 162

The E.m.f of a Cell is equals to

A

Emf(oxidation) - Emf(Reduction)

B

Emf(oxidation) + Emf(Reduction)

C

Emf(oxidation) x Emf(Reduction)

D

None of these

Q 163

Which of the following is correct relationship between Boiling point and Intermolecular forces ?

A

Boiling point increases if intermolecular forces increases

B

Boiling point decreases if intermolecular forces increases

C

Boiling point increases if intermolecular forces decreases

D

Boiling point is not affected by intermolecular forces

Q 164

Addition of 2% gypsum in cement

A

Triggers hydration

B

Triggers hydrolysis

C

Prevents rapid hardening

D

All of the above

Q 165

The increase in size of an anion is due to increase in repulsion of

A

electron-proton

B

electron-electron

C

electron-nucleus

D

proton-nucleus

Q 166

Charles's law is not being obeyed when temperature is measured on Celsius scale. That's why new scale called _____ has been developed

A

zero Fahrenheit

B

zero Kelvin

C

absolute Fahrenheit

D

all of these

Q 167

Reaction of Sodium with water is an example of

A

Reversible reaction

B

Endothermic

C

Irreversible

D

Slow

Q 168

The second bond in ethyne is:

A

Pi bond between hybridized orbitals

B

Sigma bond between hybridized orbitals

C

Pi bond between unhybridized orbitals

D

Sigma bond between unhybridized orbitals

Q 169

In Which of the following reaction of carboxylic acids only C=O group involved in bonding and OH group is replaced ?

A

Reaction of carboxylic acid with NaOH

B

Reaction of carboxylic acids with Carbonate

C

Formation of anhydride

D

Reaction of carboxylic acids with SOCl_2

Q 170

Iodoform test is performed to distinguish between

A

alcohols and phenols

B

methanol and ethanol

C

primary and secondary alcohols

D

phenols and ethers

Q 171

The molecules of CO₂ in dry ice form

A

covalent crystals

B

ionic crystals

C

metallic crystals

D

molecular crystals

Q 172

which of the following has neither secondary nor tertiary hydrogen

A

isobutane

B

pentane

C

neo-pentane

D

isopentane

Q 173

A real gas obeying Van der Waals equation will resemble to ideal gas if both a & b are

A

small

B

large

C

equal

D

none of these

Q 174

Which of the following is formed during SN_1 reactions?

A

Secondary carbocation

B

Primary carbocation

C

Tertiary carbocation

D

Methyl carbocation

Q 175

If 9.8 g of sulfuric acid dissolved in excess quantity of water, it will yield _____ moles of hydrogen ion (H^+) and _____ mole of sulphate ions (SO_4^{2-})

A

0.1, 0.2

B

0.1, 0.3

C

0.2, 0.4

D

0.2, 0.1

Q 176

The transition elements belongs to Group VIB are

A

Zn, Cd, Hg

B

Fe, Ru, Os

C

Mn, Te, Re

D

Cr, Mo, W

Q 177

Ice occupies more space than liquid water by _____?

A

0.1

0.09	B
0.05	C
0.06	D
Q 178	
If K_p and K_c have same values Δn will be	
Maximum	A
Minimum	B
Zero	C
Negligible	D
Q 179	
The element caesium bears resemblance with	
Ca	A
Cr	B
Both of the above	C
None of the above	D
Q 180	
Enzymes which bring about exchange of functional group is called_____.	
Oxidoreductase	A
Hydrolases	B
Ligases	C
	D

Transferases

English

Q 181

_____ it was raining, he went out without a raincoat.

A

Even

B

Since

C

Unless

D

Although

Q 182

I found this bangle (A) while digging (B) in the backyard.(C) I don't know who it belonged to. (D)

A

I found this bangle

B

while digging

C

in the back yard.

D

I don't know who it belonged to.

Q 183

We will need _____ umbrella today.

A

a

B

an

C

the

D

no article

Q 184

Michael _____ absent yesterday.

A

is

B

was

C

were

D

being

Q 185

Choose the correct sentence.

A

Does your train leave before eight o'clock?

B

Does your train leave before eight o'clock?

C

does your train leave before eight o'clock

D

does your train leave before eight o'clock?

Q 186

Choose the correct sentence.

A

After we'd closed the door we sat down to read the newspaper.

B

After we'd closed the door, we sat down to read the newspaper.

C

After w'ed closed the door, we sat down to read the newspaper.

D

After we'd closed the door, we sat down to read the Newspaper.

Q 187

You cannot eat that mango. It is not _____ yet.

A

best

B

pale

C

ripe

D

mature

Q 188

When I go swimming. I have to keep my eyes closed underwater.

A

complex

B

simple

C

compound

D

None

Q 189

Which one is future tense?

A

I will pay for you

B

I am coming

C

They are going

D

They were present

Q 190

There was _____ article about pollution in _____ paper.

A

a..an

B

an... the

C

a... the

D

a...a

Q 191

Sarah is so lazy. She never _____ to help her mum.

A

try

B

trying

C

tries

D

trys

Q 192

The thief _____ before the policeman reaches.

A

will have escape

B

will have escaped

C

will escaped

D

will escape

Q 193

she is

A

shes

B

shes'

C

she's

D

sh'es

Q 194

My father _____ (buy) us popcorn and orange juice.

A

buy

B

will buy

C

bought

D

have bought

Q 195

Choose the correct sentence.

A

“Trust me, I know what I’m doing,” Bilal said. “Your car is not hard to fix.”

B

“Trust me, I know what I’m doing,” Bilal said. “your car is not hard to fix.”

C

“Trust me I know what Im doing,” Bilal said. “Your car is not hard to fix.”

D

“Trust me, I know what I’m doing,” Bilal said. Your car is not hard to fix.

Q 196

He was dismissed from the service for his insistent

A

alertness

B

nobility

C

disobedience

D

alliance

Q 197

Choose the present indefinite tense form of the sentence. “He had been sleeping.”

A

He has been sleeping.

B

He had slept.

C

He has slept.

D

He sleeps.

Q 198

We have _____ small house in _____ village in _____ Netherlands.

A

the... an...a

B

an...a... the

C

a...a....the

D

a...an... the

Q 199

None of them _____ able to solve this question.

A

is

B

were

C

was

D

would be

Q 200

Find the error?

A

It has not rained since April.

B

The jurors walked solemnly into the room.

C

Had we known, we would not have come.

D

No mistakes

Logical Reasoning

Q 201

Statement The ratio is poverty is at alarming point in our country. I. The Government needs to take step for economic and development growth. II. The lower class area of people in our country needs to be supported as most of them lives hand to mouth.

A

Both of them follows

B

None of them follows

C

Only I follows

D

Only b follows

Q 202

Statement: Large number of people living in the low-lying areas has been evacuated during the last few days to safer places. The government has rushed in relief supplies to the people living in the affected areas.

A

Statement I is the cause and statement II is its effect.

B

Statement II is the cause and statement I is its effect

C

Both the statements I and II are independent causes

D

Both the statements I and II are effects of some common cause

Q 203

Complete the series AB, EF, _____, MN

A

IJ

B

RS

C

WX

D

MN

Q 204

Which word is the odd man out?

A

trivial

B

unimportant

C

important

D

insignificant

Q 205

I. The new program has been introduced with ordeal. II. The government has announced a plan on the creation of the TV programs.

A

Statement I is the cause and statement II is its effect.

B

Statement II is the cause and statement I is its effect.

C

Both statements I and II are independent causes

D

Both statements I and II are the effects of independent cause.

Q 206

Gym is to exertion as consuming is to

A

Food

B

Dieting

C

Fitness

D

Eatery

Q 207

Arrange the following words in a meaningful sequence. 1.Infection 2.Consultation 3.Doctor 4.Treatment 5.Recovery

A

1, 3, 4, 5, 2

B

1, 3, 2, 4, 5

C

1, 2, 3, 4, 5

D

2, 3, 5, 1, 4

Q 208

Statement The availability of imported fruits has increased in the indigenous market and so the demand for indigenous fruits has been decreased. I. To help the indigenous producers of fruits, the Government should impose high import duty on these fruits, even if these are not of good quality. II. The fruit vendors should stop selling imported fruits. So that the demand for indigenous fruits would be increased.

A

both of them follows

B

None of them follows

C

Only I follows

D

Only II follows

Q 209

Statement: There has been a high increase in the incidents of atrocities against women in the city during the past few months. The police authority has been unable to nab the culprits who are committing crime against women.

A

Statement I is the cause and statement II is its effect.

B

Statement II is the cause and statement I is its effect

C

Both the statements I and II are independent causes

D

Both the statements I and II are effects of independent causes

Q 210

Fact 1 Pictures can tell a story Fact 2 All storybooks have a picture Fact 3 Some story books have words If the above three statements are facts than which of the following statement will also be a fact I. Pictures can tell a story better than words

can II. The stories in storybooks are simple III. Some story books have both pictures and words

A

Only I

B

Only II

C

Only III

D

None of them is a fact

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